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First-principles-derived kinetics of the reactions involved in low-temperature dimethyl ether oxidation

A. Andersen and E.A. Carter

In Table 1, column '$k_\infty(T)$', the high-pressure rate expression for the reaction of Equation (10) should read: $3.24 \times 10^{12} T^{0.31} e^{-23.7/RT}$.

In Table 3, column 'k_∞(expression)', the rate expression for the reaction of Equation –(6) should read: $3.59 \times 10^{19} T^{-1.03} e^{-29.5/RT}$.

In Table 6, the ΔH°_f for CH_2OO should be $+24.7\,\text{kcal mol}^{-1}$ (this work).

Reference (66), was published in 2006:
A. Andersen and E.A. Carter, J. Phys. Chem. A **110**, 1393 (2006).

The high-pressure rate constant from Maricq *et al*. [39] in the second column of page 372 should read: $1.87 \times 10^{12} \exp(+0.6\,\text{kcal mol}^{-1}/RT)\,\text{cm}^3\,\text{mol}^{-1}\,\text{s}^{-1}$.

The high-pressure rate constant from Curran *et al*. [4,5] in the first paragraph of the left column on page 377 should read: $1.50 \times 10^{13} \exp(-20.8\,\text{kcal mol}^{-1}/RT)\,\text{s}^{-1}$.

Taylor and Francis wish to apologise for any inconvenience caused.

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